Sulfur isotopic measurements from a West Antarctic ice core: implications for sulfate source and transport

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ABSTRACT. Measurements of $\delta^{34}S$ covering the years 1935–76 and including the 1963 Agung (Indonesia) eruption were made on a West Antarctic firm core, RIDSA (78.73 S, 116.33 W; 1740 m a.s.l.), and results are used to unravel potential source functions in the sulfur cycle over West Antarctica. The $\delta^{34}S$ values of SO$_2^{\text{aq}}$ range from 3.1% to 9.9%. These values are lower than those reported for central Antarctica, from near South Pole station, of 9.3–18.1% (Patris and others, 2000). While the Agung period is isotopically distinct at South Pole, it is not in the RIDSA dataset, suggesting differences in the source associations for the sulfur cycle between these two regions. Given the relatively large input of marine aerosols at RIDSA (determined from Na$^+$ data and the seasonal SO$_2^-$ cycle), there is likely a large marine biogenic SO$_2^-$ influence. The $\delta^{34}S$ values indicate, however, that this marine biogenic SO$_2^-$, with a well-established $\delta^{34}S$ of 18‰, is mixing with SO$_2^-$ that has extremely negative $\delta^{34}S$ values to produce the measured isotope values in the RIDSA core. We suggest that the transport and deposition of stratospheric SO$_2^-$ in West Antarctica, combined with local volcanic input, accounts for the observed variance in $\delta^{34}S$ values.

INTRODUCTION

Deconvoluting the biogeochemical cycling of atmospheric sulfur is important for understanding climate change. Sulfate (SO$_2^-$) aerosols play a key role in moderating the Earth’s climate. They function as cloud condensation nuclei and can enhance the lifetime of clouds and their ability to reflect incoming solar radiation. Additional aerosol mass in the atmosphere results in an increase in optical thickness and cloud albedo (Anderson and others, 1995). This effect could in turn cause a change in oceanic productivity (which is a source of SO$_2^-$ aerosols), thereby creating a climate feedback loop that is susceptible to climate change (Bates and others, 1987; Charlson and others, 1987). Sulfate aerosols also contribute to the formation of SO$_2^-$ haze. The size of SO$_2^-$ particles determines their effect on climate. Sulfate particles derived from sea-salt particles are generally >1 μm in diameter and play an important role in the marine boundary layer (Murphy and others, 1998). Sulfate aerosols play various other roles in atmospheric chemistry that could be affected by or affect climate change, including an alteration of oxidation processes in the marine boundary layer (Andreae and Crutzen, 1997).

The Antarctic displays high sensitivity to climate change and in turn greatly impacts the climate of low latitudes (Bromwich and Parish, 1998). One of the limitations to understanding the sulfur cycle is the assessment of how much each potential SO$_2^-$ source contributes and the accompanying effect on atmospheric processes and climate. There are three recognized sources of SO$_2^-$ in Antarctica: sea-salt, marine biogenic and volcanic (Delmas, 1982; Legrand, 1997). Anthropogenic source contributions to Antarctica are believed to be negligible (Shaw, 1982; Legrand and Mayewski, 1997). Unlike ice-core records from Greenland, which clearly show an increase in background SO$_2^-$ concentration since the industrial revolution (Mayewski and others, 1993), Antarctica shows no increase in total SO$_2^-$ concentration with time (Legrand, 1997; Legrand and Mayewski, 1997). The relatively few sources of SO$_2^-$, when compared to Greenland, make Antarctica a useful environment to observe the natural sulfur cycle.

Estimates of the Antarctic SO$_2^-$ budget are based on aerosol measurements, analysis of snow and ice samples, and an understanding of volcanism and meteorology of the region. Aerosol studies indicate that background xsSO$_2^-$ in the Antarctic comes primarily from marine biogenic sources (Pszenny and others, 1989; Minikin and others, 1998). Mount Erebus, the only active volcano on the continent, contributes to the Antarctic sulfur cycle, but the strength of this source varies temporally depending on the size of the lava lake (Kyle and Meekler, 1990).

Ice-core records provide an historical record of atmospheric chemistry, including SO$_2^-$ deposition (Delmas and Boutron, 1977; Delmas, 1982; Legrand and others, 1991; Legrand, 1997) and have been used to assess the sulfur cycle in Antarctica (Delmas and Boutron, 1980; Legrand and Feniet-Saigne, 1991; Meyerson and others, 2002). Snow pits and ice cores demonstrate that the SO$_2^-$ production operates seasonally, with concentration minima in the winter and maxima in the summer (Legrand and Pasteur, 1998; Minikin and others, 1998). Excess SO$_2^-$ concentrations are independent of accumulation rate and elevation on a continent-wide scale (Mulvaney and Wolff, 1994; Kreutz and Mayewski, 1999). Distance inland correlates to a decrease in the concentration of xsSO$_2^-$, which may be due to the portion of SO$_2^-$ that is derived from marine productivity. The 1998 Italian ITASE (International Trans-Antarctic Scientific Expedition) traverse in East Antarctica demonstrated that the decrease inland of xsSO$_2^-$ extended only 250 km and then SO$_2^-$ increased in concentration from 250 to 770 km (Propisito and others, 2002). In surface snow measurements, there is a negative correlation between both distance inland and elevation and methanesulfonic acid...
(MSA) concentration. As with $\text{xsSO}_2^-$, accumulation rate does not affect MSA concentration, except perhaps at low-accumulation-rate sites (Legrand and others, 1992; Wagnon and others, 1999). A seasonal variation of MSA occurs in several cores. The presence of the seasonal signal depends on location, with only an interannual signal at South Pole and a stronger signal in coastal areas (Legrand and others, 1992; Meyerson and others, 2002). In other Antarctic studies, there is no seasonal variation, unlike both sea-salt $\text{SO}_2^-$ and $\text{xsSO}_2^-$ (Ivey and others, 1986). Recent aerosol studies and data from snow pits and firm cores indicate that MSA and $\text{xsSO}_2^-$ correlate on a seasonal basis, with summer peaks and winter lows (Legrand and others, 1992; Stenni and others, 2000; Arimoto and others, 2001).

The $\text{SO}_2^-$ sources and emission rates are important indicators of the vertical distribution of $\text{SO}_2^-$ aerosols over Antarctica, and therefore the radiative effects. Stable sulfur isotopes provide a potentially useful tool for estimating $\text{SO}_2^-$ sources in ice cores. This technique has previously been used to determine the source of $\text{SO}_2^-$ in precipitation and aerosol studies (Callhoun and others, 1991; Nriagu and others, 1991; Wadleigh and others, 1996). Sulfur isotopes are useful in studying atmospheric transport mechanisms and chemistry during volcanic eruptions (Castleman and others, 1973) and during anthropogenic emissions (Newman and others, 1975). Differences in chemical reaction rates can cause isotope fractionation, resulting in an isotopic ‘signature’ that can be used to identify the chemical process (such as incorporation into a biologic system) or source of the atmospheric $\text{SO}_2^-$. The isotopic signature of sulfur can be used to indicate the source of the $\text{SO}_2^-$ if there are a few sources with distinct signatures, as is the case in Antarctica (Nielsen, 1974; McArdle and Liss, 1995). The resulting $\delta^{34}S$ can be deconstructed through a mixing equation. The more that is understood about the sources of atmospheric sulfur in a region, the easier it becomes to use sulfur isotopes to quantify $\text{SO}_2^-$ sources.

Recently, sulfur isotopes were used in an ice-core study in central Antarctica (Patris and others, 2000). A quantitative assessment of $\text{SO}_2^-$ at South Pole was made based on sulfur isotopic measurements on sections of a firm core. Patris and others (2000) determined that marine biogenic emissions are the dominant background source of $\text{SO}_2^-$. Sulfur isotopes in a Greenland ice core were also used to examine the Arctic sulfur cycle during both pre-industrial and industrial times (Patris and others, 2002). Applying sulfur isotope studies to ice-core records can greatly improve understanding of the sulfur cycle over long periods of time by providing evidence for $\text{SO}_2^-$ source regions and transport.

The $\delta^{34}S$ measurements presented here are the first sulfur isotopic measurements from the West Antarctic ice sheet. These measurements are used to infer differences between the sources of $\text{SO}_2^-$ in West Antarctica and central Antarctica. These data indicate the potential of sulfur isotopes as a tool for partitioning the $\text{SO}_2^-$ sources (marine biogenic, sea-salt and volcanic) to West Antarctica.

**METHODS**

**Core processing**

The 147 m RIDSA ice core (78.73 S, 116.33 W; 1740 m a.s.l.) (Fig. 1) was retrieved in West Antarctica in 1995 (Kreutz and Mayewski, 1999). Core processing was performed in clean conditions, with samples collected into pre-cleaned containers and stored below −150°C until melting and analysis. Processing and chemical analysis were done at 3 cm intervals down to 60 m, yielding high-resolution (sub-annual) chemical data. From meter 60 to 147, processing was done at 60 cm intervals. One-meter sections of core 10 to 26 m were used for sulfur isotopic analysis, covering the period 1935–76. The time covered by individual samples in this dataset (RIDSA 2001) was 2–3 years, depending on accumulation rate.

Sections of the RIDSA core from 67.2 to 69.6 m, covering the years 1976–2005, and from 13.8 to 15.9 m, covering the years 1961–67, were used for isotopic method development. These samples are designated as RIDSA 2000.

**Ion analysis**

Major-ions (Na$^+$, K$^+$, NH$_4^+$, Mg$^{2+}$, Ca$^{2+}$, Cl$^-$, NO$_3^-$, SO$_4^{2-}$) and MS (measured as methanesulfonic acid or MSA) measurements were made by ion chromatography at the (parts per billion) ppb level. Anions were analyzed on a Dionex DX-500 ion chromatograph with an AS-11 column using 6 mM NaOH eluent. Cations were analyzed on a Dionex DX-500 ion chromatograph with a CS-12A column using 25 mM MSA eluent. A β-activity profile was made using 20 cm samples (Kreutz and others, 2000). The β-activity maximum is assumed to indicate the austral summer 1964/65, based on the global peak from the Atmospheric Test Ban Treaty (1963) and the transport lag to Antarctica (Picciotto and Wilgain, 1963). It is possible to count back to the global horizon using summer peaks in nssSO$_4^{2-}$. Dating is therefore based on seasonal nssSO$_4^{2-}$ cycles (Kreutz and others, 2000). These dating techniques show that the entire core covers the time period 1506–1995.

The excess SO$_4^{2-}$ is calculated from the Na$^+$ concentration of the samples using the following equation:
Table 1. Ion concentrations for each RIDSA sample. Units are ng g\(^{-1}\) except for Cl/Na mass ratio and R value, which have no units. The R value is the ratio MSA/Na\(^{2+}\). is and is often used to assess the marine biogenic portion of total SO\(_4^{2-}\).

<table>
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<tr>
<th>Depth (m)</th>
<th>Na(^{+})</th>
<th>NH(_4)(^{+})</th>
<th>K(^{+})</th>
<th>Mg(^{2+})</th>
<th>Ca(^{2+})</th>
<th>Cl(^{-})</th>
<th>NO(_3^{-})</th>
<th>MSA</th>
<th>NSS SO(_4^{2-})</th>
<th>Cl(^{-})/Na(^{+}) mass ratio</th>
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\[
\text{SO}_4^{2-} = \left[ \text{SO}_4^{2-} \right] - k[\text{Na}^+] \tag{1}
\]

where \( k \) is equal to the sea-water (\( \text{SO}_4^{2-}/\text{Na}^+ \)) mass ratio, 0.251 (Wilson, 1975).

**Isotope analysis**

One hundred micrograms of SO\(_4^{2-}\) is necessary for isotopic analysis, making high-resolution analysis difficult in Antarctic cores due to low accumulation rates (Patris and others, 2000). The RIDSA 2000 samples are aggregated residual from samples taken from the RIDSA core that were previously aliquoted for chemical analysis. Sampling was initially performed at 3 cm intervals for analysis of major ions. Two time periods, one covering the Agung (Indonesia) period (1963 eruption) and one covering a pre-industrial period, were chosen, and sample remaining from the initial chemical analysis was combined until a volume of 1 L was attained. The years covered by these two measurements were dependent on the volume needed for isotopic analysis (1 L). Additional isotopic analysis was performed on 1 m sections of the RIDSA core, including the Agung period. All core sections were melted in sterile conditions and then transferred to 1 L bottles for analysis at the Memorial University of Newfoundland. Sample aliquots for stable-isotopic analysis were transferred to acid-washed ‘snap-cap’ vials. Vials were placed on a hot plate in a positive-pressure, High Efficiency Particle Air (HEPA)-filtered air fume hood, and the solutions were slowly evaporated until the volume was reduced to approximately 20 mol from an initial 1 L. The concentrated sample was then used to fill 10 to 10 to 10 mM tin capsules and evaporated to dryness. The process was repeated between two and five times according to the measured SO\(_4^{2-}\) concentration in order to obtain sufficient sample for isotopic analysis. After the final evaporation, the tin capsule was closed and placed in the auto-sampler of a Carlo Erba 1500 elemental analyzer interfaced by ConFlo II to a Finnigan MAT 252 stable-isotope ratio mass spectrometer. The sulfur isotopic ratio is expressed in delta notation with respect to the standard, Vienna Canyon Diablo Troilite (VCDT), and is represented by:

\[
\delta^{34}S(\%) = \left\{ \frac{\left[ \frac{34}{32}S \right]_{\text{sample}}}{\left[ \frac{34}{32}S \right]_{\text{VCDT}}} - 1 \right\} \times 1000. \tag{2}
\]

Calibration to VCDT was performed using two Newfoundland internal standards, NZ-1 and NZ-2, with reported values of -0.30 ± 0.3% and +21.0 ± 0.3% respectively. The overall error on isotopic analysis based on duplicates and internal standards is ±0.5%.

**RESULTS**

**Ion data**

Ion averages for each of the meters taken for isotopic analysis (10–26 m; RIDSA 2001) and the RIDSA 2000 samples are presented in Table 1, along with the years covered. One sample of overlap between the two RIDSA datasets is consistent with the other.

Sodium and xsSO\(_4^{2-}\) have strong seasonal cycles, consistent with other studies in the Antarctic (Fig. 2) (Murozumi and others, 1969; Herron and Langway, 1979; Hammer, 1980; Legrand and others, 1991; Whitlow and others, 1992; Minikin and others, 1998; Kreutz and Mayewski, 1999; Reusch and others, 1999).

Na\(^{+}\) peaks in the winter due to increased storm activity, while xsSO\(_4^{2-}\) has a summer peak due to its marine biogenic component (Whitlow and others, 1992). During the summer months, sea ice retreats as solar radiation increases, leaving more open ocean in which primary productivity can thrive. One sign of increased productivity might be the increase in SO\(_4^{2-}\) concentration during the summer months. The MSA
profile shows a subdued seasonal signal relative to the \( \text{xsSO}_4^{2-} \) signal (Fig. 2). In other glaciochemical studies performed in the Antarctic, MSA has also been found to correlate with \( \text{xsSO}_4^{2-} \), with both showing a summer maximum (Kreutz and Mayewsky, 1999). The \( \text{xsSO}_4^{2-} \) is a combination of primarily, or perhaps exclusively, marine biogenic and volcanic \( \text{SO}_4^{2-} \). Table 2 gives concentration data for the three possible contributors to total \( \text{SO}_4^{2-} \) for the RIDSA samples.

The Agung eruption (17 March 1963) is apparent in the \( \text{xsSO}_4^{2-} \) profile (Fig. 2) and can be seen as an increase in the seasonal minima of \( \text{SO}_4^{2-} \) from late 1963 to 1966 (Devine and others, 1984). The eruption is recorded in the RIDSA core starting in late 1963, about 6 months after the eruption. This time difference of around 6 months between the eruption and its subsequent deposition in the Antarctic is due to the long distance the \( \text{SO}_4^{2-} \) load is transported. The lag time varies depending on the location of the volcano, but other work done in the Antarctic shows a similar 6 month time lag for the Agung eruption (Legrand, 1997). Other Antarctic cores show volcanic peaks as a \( \text{SO}_4^{2-} \) spike above all other seasonal maxima. In the RIDSA core, however, volcanic input is typically represented by an increase in background \( \text{SO}_4^{2-} \), as is the case with the Agung eruption (Fig. 2). The other volcanic period shown in Figure 2, the 1991 Pinatubo–Hudson eruption, is an exception. This may be related to seasonal changes in atmospheric transport or accumulation rate.

Isotope data
The \( \delta^{34}S \) of each meter of core analyzed is given in Table 2. The values range from 3.1‰ to 9.9‰, which is lower than those reported for central Antarctica, which range from 9.3‰ to 18.1‰ (Patris and others, 2000). The measured isotopic signature is plotted against the inverse of the total \( \text{SO}_4^{2-} \) concentration to highlight isotope trends with concentration (Fig. 3).

The East Antarctic data show a clear trend towards lower \( \delta^{34}S \) values with increased concentration, and a linear regression (\( r=0.997 \)) yields a value of 2.6‰ for the intercept, which is assumed to be predominantly volcanic \( \text{SO}_4^{2-} \).

In West Antarctica, there is no discernible trend with \( \text{SO}_4^{2-} \) concentration for the raw data. The Agung period, which was isotopically distinct in the East Antarctic data, is not distinct in West Antarctica. The variability, or scatter, in the RIDSA data may be due in part to the sampling technique. The core was sampled continuously at 1 m intervals. This resulted in ten samples containing two summer sections, and six samples containing three summer sections.

The isotopic signature of the samples can be broken down into its components (Patris and others, 2000):

\[
\delta_{\text{mb}} = f_{\text{mb}} \delta_{\text{mb}} + f_{\text{vol}} \delta_{\text{vol}} + f_{\text{sub}} \delta_{\text{sub}},
\]

(3)

where \( \delta_{\text{mb}} \) is the measured isotopic signature as given in Table 2, \( \delta_{\text{vol}} \) and \( \delta_{\text{sub}} \) are the isotopic signatures of the sea-salt, volcanic and marine biogenic components of the sample, and \( f_{\text{mb}} \), \( f_{\text{vol}} \) and \( f_{\text{sub}} \) are the mass fractions of the sea-salt, excess, volcanic and marine biogenic components, respectively. \( f_{\text{mb}} \) and \( f_{\text{vol}} \) can be determined using the concentration for \( \text{xsSO}_4^{2-} \), which was computed using the \( Na^{+} \) concentration. The following equation represents the sea-salt fraction of a sample:

\[
f_{\text{mb}} = \left( \frac{[\text{SO}_4^{2-}]_{\text{mb}} - [\text{xsSO}_4^{2-}]}{[\text{SO}_4^{2-}]} \right).
\]

(4)

The non-sea-salt fraction is \( (1-f_{\text{mb}}) \). The isotopic signature of sea-salt \( \text{SO}_4^{2-} \) is well constrained through field study to be 21‰ (Rees and others, 1978). This reflects an enrichment in \( ^{34}S \) in ocean water compared to terrestrial sulfur. Rearranging Equation (3) yields:

\[
\delta_{\text{ss}} = (\delta_{\text{mb}} - f_{\text{mb}} \delta_{\text{sub}})/f_{\text{ss}},
\]

(5)

Substituting the established value for \( \delta_{\text{ss}} \) into the equation,
along with the calculated $f_{\text{so}}$ and $f_{\text{ss}}$, yields an isotopic signature for the excess component of the samples (Table 2). These values are plotted against the inverse of the $x_{\text{SSO}_4^{2-}}$ concentration to yield a mixing diagram for the $x$ component in Figure 4. No significant correlation exists ($r = 0.38$). The highest $\delta^{34}$S value (in the RIDSA time series) occurs in the sample covering the years 1947–50 (Fig. 5a). Mount Erebus eruptions are noted on the time series. The original $x_{\text{SSO}_4^{2-}}$ data were smoothed using a high-tension robust spline to produce the data in Figure 5b. The robust-spline technique differs from other statistical procedures because it is more resistant to outliers, providing a smoothing technique that yields useful information without underestimating sporadic processes (Meeker and others, 1995). The peaks in the smoothed data reflect changes in volcanic $\text{SO}_4^{2-}$. There appears to be a relationship between increases in background $x_{\text{SSO}_4^{2-}}$ and $\delta^{34}$S.

### DISCUSSION

The sulfur isotope ratios in the RIDSA ice core are lower than those in central Antarctica (Patris and others, 2000). Caution must be used when looking at sulfur isotopic signatures as a source fingerprint because while processes such as atmospheric transport generally lead to similar trends, the $\delta^{34}$S value can still vary greatly for any different source (Castleman and others, 1973). For example, $\delta^{34}$S values from volcanic emissions have been estimated to range from $-15\%$ to $+25\%$ (McArthur and Liss, 1999) and from $-5\%$ to $5\%$ (Nielsen, 1974). Taking this uncertainty into account, there are several explanations for the West Antarctic sulfur isotope data. Because the $\delta^{34}$S values from West Antarctic fall within the range of newly erupted volcanic material, the background $\text{SO}_4^{2-}$ source might at first appear to be entirely of volcanic origin. The $\delta^{34}$S values of sulfur from newly erupted basaltic range from $2\%$ to $7\%$ (De Hoog and others, 2001). Studies of volcanic gas emissions show a positive $\delta^{34}$S trend (Castleman and others, 1974). There was active volcanism in Antarctica during the time period of the RIDSA core. The Smithsonian database of volcanic eruptions (www.volcano.si.edu/vgp) lists Mount Erebus, Deception Island, Candlemas and Bristol Island as known eruptive volcanoes between 1935 and 1976. It is unlikely that the low $\delta^{34}$S values can be attributed solely to local volcanism. It is clear that there is a strong marine influence on the RIDSA site. Sea-salt $\text{SO}_4^{2-}$ makes up 11.7–35.3% of the total $\text{SO}_4^{2-}$ concentration for the samples (Table 2). The isotopic signature of $\text{SO}_4^{2-}$ from a marine biogenic source is reasonably well constrained at $-18\%$ (Calhoun and others, 1991; Patris and others, 2000). This value is slightly lower than that of sea-salt $\text{SO}_4^{2-}$, which is close to 21% (Rees and others, 1978). The difference between these two marine constituents is the result of fractionation that occurs in the biological use of sea-water $\text{SO}_4^{2-}$. Because there is a strong sea-salt component to the total $\text{SO}_4^{2-}$ at the RIDSA site, it is reasonable to assume that there is a strong marine biogenic component. An isotopic study of the sulfur cycle in central Antarctica determined that the background source of sulfur was almost exclusively of marine biogenic origin (Patris and others, 2000). In West Antarctica, however, the background isotopic signature is too depleted in $^{34}$S to be of marine origin. The $\delta^{34}$S values for the $x$ component of the RIDSA samples range from $-0.7\%$ to $6.8\%$. Factoring out the sea-salt contribution reduces the $\delta^{34}$S values, suggesting that the $\delta^{34}$S values from $x_{\text{SSO}_4^{2-}}$ represent some combination of marine biogenic and volcanic $\text{SO}_4^{2-}$. The heavier isotope ($^{32}$S) is depleted faster than the lighter isotope ($^{34}$S) after volcanic eruptions, due to changes in

### Table 2. Sulfur data from the RIDSA core. $f_{\text{so}}$ indicates the fraction of the total $\text{SO}_4^{2-}$ concentration that comes from sea salt, $\delta_{\text{so}}$ is the isotopic signature for total $\text{SO}_4^{2-}$ and $\delta_{\text{ss}}$ is the isotopic signature for the non-sea-salt portion of total $\text{SO}_4^{2-}$.

<table>
<thead>
<tr>
<th>Depth (m)</th>
<th>Years</th>
<th>$[\text{SO}_4^{2-}]$ (ng g$^{-1}$)</th>
<th>Sea salt %</th>
<th>Biogenic + volcanic %</th>
<th>$f_{\text{so}}$</th>
<th>$\delta_{\text{so}}$</th>
<th>$\delta_{\text{ss}}$</th>
</tr>
</thead>
<tbody>
<tr>
<td>RIDSA 2001</td>
<td>10–11</td>
<td>1976–74</td>
<td>27.9</td>
<td>5.1</td>
<td>22.7</td>
<td>0.18</td>
<td>4.5</td>
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<td></td>
<td>11–12</td>
<td>1974–71</td>
<td>34.2</td>
<td>5.7</td>
<td>28.4</td>
<td>0.17</td>
<td>5.5</td>
</tr>
<tr>
<td></td>
<td>12–13</td>
<td>1971–69</td>
<td>29.5</td>
<td>6.5</td>
<td>22.9</td>
<td>0.22</td>
<td>5.1</td>
</tr>
<tr>
<td></td>
<td>13–14</td>
<td>1969–66</td>
<td>36.6</td>
<td>6.6</td>
<td>29.8</td>
<td>0.18</td>
<td>4.8</td>
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<tr>
<td></td>
<td>14–15</td>
<td>1966–63</td>
<td>56.2</td>
<td>6.6</td>
<td>49.5</td>
<td>0.12</td>
<td>5.4</td>
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<tr>
<td></td>
<td>15–16</td>
<td>1963–60</td>
<td>36.6</td>
<td>5.0</td>
<td>31.5</td>
<td>0.14</td>
<td>4.2</td>
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<tr>
<td></td>
<td>16–17</td>
<td>1960–58</td>
<td>26.2</td>
<td>5.3</td>
<td>20.8</td>
<td>0.20</td>
<td>5.6</td>
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<td>1958–55</td>
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<td>4.3</td>
<td>24.8</td>
<td>0.15</td>
<td>6.7</td>
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<td>18–19</td>
<td>1955–53</td>
<td>38.5</td>
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<td>28.5</td>
<td>0.26</td>
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<td></td>
<td>19–20</td>
<td>1953–50</td>
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<td>29.4</td>
<td>0.16</td>
<td>6.0</td>
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<td>20–21</td>
<td>1950–47</td>
<td>35.1</td>
<td>7.6</td>
<td>27.4</td>
<td>0.22</td>
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<td>21–22</td>
<td>1947–45</td>
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<td>4.3</td>
<td>32.3</td>
<td>0.12</td>
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<td>22–23</td>
<td>1945–43</td>
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<td>4.8</td>
<td>22.4</td>
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<td>1943–40</td>
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<td>6.0</td>
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<td>0.18</td>
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<td>24–25</td>
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<td>4.4</td>
<td>22.3</td>
<td>0.16</td>
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<td>25–26</td>
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<td>9.6</td>
<td>17.3</td>
<td>0.35</td>
<td>7.5</td>
</tr>
<tr>
<td>RIDSA 2000</td>
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<td>1961–67</td>
<td>42.9</td>
<td>6.5</td>
<td>36.3</td>
<td>0.15</td>
<td>4.8</td>
</tr>
<tr>
<td></td>
<td>67.2–69.6</td>
<td>1796–1805</td>
<td>33.1</td>
<td>7.2</td>
<td>25.9</td>
<td>0.22</td>
<td>7.2</td>
</tr>
</tbody>
</table>

The calculations are explained in the text.
stratospheric chemistry and the pathway of $\text{SO}_2^{2-}$ aerosol formation. Plume studies from anthropogenic emissions confirm this mechanism (Castleman and others, 1974; Newman and others, 1975). The lowest $\delta^{34}$S values (–24.4‰) occur higher in the stratosphere (Castleman and others, 1974). One possibility in West Antarctica is that a large portion of the sulfur budget comes from a marine biogenic source with a $\delta^{34}$S value around 18‰, and that the rest of the $\text{SO}_2^{2-}$ can be attributed to aged and fractionated stratospheric air. Stratospheric sulfur with extremely low $\delta^{34}$S values could mix with local air masses with a relatively high $\delta^{34}$S value to produce the intermediate values (Fig. 4). For example, using a 50:50 input from stratospherically transported materials and marine materials (18‰), a calculation can be made from the overall $\delta^{34}$S to determine the $\delta^{34}$S of the second contributor. Using the following mixing relationship between recalculated $\delta^{34}$S value and its two components

$$\delta_{new} = f_{abh}\delta_{abh} + f_{vol}\delta_{vol}$$  (6)

and substituting an $\delta^{34}$S value of 2‰, within the range of the recalculated $\delta^{34}$S values, gives the $\delta^{34}$S value of the unknown contributor as –14‰. If stratospheric transport and temperature effects yield $\delta^{34}$S values that are even more negative than –14‰, then the ratio of marine biogenic to stratospheric input would be higher, indicating that local tropospheric deposition is of greater consequence to the West Antarctic sulfur cycle than stratospheric input.

The variance within the time series (–0.7‰ to 6.8‰; Fig. 5) corresponds to changes in the smoothed background $\text{SO}_2^{2-}$ signal (Fig. 5). Increases in the $\delta^{34}$S values over time correspond to increases in the $\text{SO}_2^{2-}$ concentration over time. These $\text{SO}_2^{2-}$ peaks do not appear to reflect any global volcanism. Mount Erebus, however, is known to have erupted during the time period covered by each of the samples that has a peak in the time series. The largest $\delta^{34}$S value is 6.8‰, in the sample covering the years 1947–50. It coincides with a peak in the background $\text{SO}_2^{2-}$ and a known Erebus eruption. Local volcanism would cause an increase in the $\delta^{34}$S value because newly erupted material has a positive $\delta^{34}$S value and is deposited relatively quickly, without the fractionation associated with transport (Castleman and others, 1974). The $\delta^{34}$S peak during the 1947 Erebus event is much larger than the peak associated with the two other local events that occurred during 1957 and 1972. The peak in $\text{SO}_2^{2-}$ during the 1947 event is also much larger than the other peaks and is further evidence of the importance of local volcanism for the West Antarctic sulfur cycle (Delmas, 1982). There are not unusually high peaks in $\delta^{34}$S associated with the 1957 and 1972 events and this may be because they were weak eruptions relative to the 1947 Erebus event and therefore did not deposit enough $\text{SO}_2^{2-}$ to impact the isotopic signature in the same way.

The $\delta^{34}$S value during the Agung period could be expected to be very depleted in the heavier ($^{34}$S) isotope and show a minimum in the dataset based on the assumption that stratospheric $\text{SO}_2^{2-}$ (which is depleted in $^{34}$S) is mixing with $\text{SO}_2^{2-}$ from a marine biogenic source. It actually shows a higher value than the samples it precedes and follows. Local volcanism (Mount Erebus) also impacted deposition during this time and may provide an explanation for this discrepancy. If the 1963 Erebus eruption caused a higher $\delta^{34}$S in local precipitation, then the subsequent mixing with precipitation from Agung (which carries a low $\delta^{34}$S) might provide a mechanism by which the overall measured $\delta^{34}$S in the sample might not be as strongly affected by stratospheric air as it would during periods which are free from global volcanic events, such as the data from the 1947–50 sample.

**CONCLUSIONS**

The first sulfur isotope measurements have been made in a West Antarctic ice core and used to assess $\text{SO}_2^{2-}$ deposition.
in West Antarctica from 1935 to 1976. The measured δ34S values were recalculated to reflect only the xs component of total SO2−2, which in Antarctica includes only marine biogenic and volcanic contributions. These xsSO2−2 δ34S values range from −0.7% to 6.8%, lower than those reported in central Antarctica (Patris and others, 2000). Based on the relatively high percentage of sea-salt SO2−2 in the RIDS A core, it can be assumed that there is also a substantial amount of marine biogenic SO2−2 in each sample. The xsδ34S values are too low to represent solely a marine biogenic influence. Volcanic plume studies have shown that volcanic emissions into the stratosphere are initially positive with respect to δ34S, and then become progressively lower over time due to a change in the chemical pathway of SO2−2 aerosol formation. δ34S values decrease with time and vertical distribution after an eruption (Castleman and others, 1973). Samples taken at the top of the stratosphere show more negative δ34S values than those lower in the stratosphere. This stratospheric air, which is extremely depleted in 34S, could be mixing with local, higher δ34S air to produce the values seen in West Antarctica. Local volcanic emissions, specifically from Mount Erebus, also play a role in the sulfur cycle in West Antarctica, based on the correlation between increased δ34S values, increased background xsSO2−2 and Mount Erebus eruptions.

Background peaks in xsSO2−2 from West Antarctica can be attributed to local volcanic input. The low resolution of each sample, however, hinders this interpretation of Antarctic SO2−2 production. High-resolution δ34S samples are needed to dissect further the seasonality of the isotopic signature. Once this is determined, it may be possible to apply the use of sulfur isotopes to ice-core studies to more accurately assess the past SO2−2 production.

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